Roll No.: _____   Enrolment No._____

PARUL UNIVERSITY

PARUL INSTITUTE OF PHARMACY

B.PHARM FIRST SEMESTER

IMPROVEMENT THEORY EXAMINATION: 2022-23

Subject Name: Pharmaceutical Analysis-I
Subject Code: BP102T   **Date: 28/02/2023**
Time: 1 hr 15 min   **Total Marks: 30**

Instructions:
1. **Make suitable assumptions wherever necessary.**
2. **Figures to the right indicate maximum marks.**

Q.1   Multiple Choice Questions. (10 X 1=10)   (10)

1	Method is not used for minimizing errors	
	a) Running a blank determination	b) Standard addition
	c) Calibration of apparatus	d) Co- precipitation
2	Precision refers as :	
	a) Logical accuracy	b) Measurement with internal standards
	c) Closeness of true value to experimental value	d) Degree of reproducibility
3	The first edition of United States of Pharmacopoeia was published in	
	a) 1860	b) 1854
	c) 1820	d) 1843
4	If 2 gm NaOH is dissolved in 500ml water –what should be the normality of solution?	
	a) 0.5 N	b) 0.2 N
	c) 0.1 N	d) 0.05 N
5	Full Form of IPC is	
	a) Indian Pharmacy community	b) Indian Pharmacological Code
	c) Indian Pharmacopoeia commission	d) Indian Pharmaceutical commitee
6	Consider the titration of 0.1N HCl Vs 0.1N NaOH. What may be the pH at equivalence point of same acid base titration?	
	a) 2	b) 8
	c) 7	d) None of the above
7	All of the following statements are incorrect for buffers excluding…………..	
	a) The pH of buffer solution is neutral	b) The pH of buffer solution is constant after addition of small amount of acid/ base
	c) The pH of buffer solution is acidic after addition of small amount of acid/ base	d) The pH of buffer solution is basic after addition of small amount of acid/ base

8	Which of the following statement is correct for strong electrolytes in acid base titration?	
	a) Strong electrolytes ionize completely in respective solvents	b) Strong electrolytes ionize partially in respective solvents
	c) The ionization of strong electrolytes will depends upon solvent properties	d) All may be considered as correct statements for strong electrolytes ionization
9	The solvents having high affinity for H^+ ions are known as…………..	
	a) Protogenic solvents	b) Protophilic solvents
	c) Amphiprotic solvents	d) Aprotic solvents
10	"The rate of chemical reaction is proportional to the active masses of the reacting substances." – The statement is represents……………	
	a) Law of equilibrium	b) Law of mass action
	c) Law of dissociation	d) Law of ionization

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Q.2   **Long Answer: (Answer Any one) (1 X 10 =10)**   **(10)**
       1) Define Buffer. Explain the specific properties of buffer solution.
       Derive the Henderson-Hasselbalch equation in detail.
       2) Define Indicator. Write a note on theories of acid base indicators.

Q.3   **Short Answer (Answer Any Two) (2 X 5=10)**   **(10)**
       1) Write about systematic error in detail.

       2) Write in detail about precision and accuracy.

       3) Enlist the techniques of Quantitative analysis and explain about
       titrimetric methods.