

Roll No.: _____

Enrolment No. _____

PARUL UNIVERSITY
SCHOOL OF PHARMACY
PILARM FIRST SEMESTER

FIRST INTERNAL THEORY EXAMINATION: 2021-22

Subject Name: Pharmaceutical Inorganic Chemistry

Subject Code: BP104T

Time: 10:00 am to 11:15am

Date: 09/12/2021

Total Marks: 30

Instructions:

1. Figures to the right indicate full marks.
2. Make suitable assumptions wherever necessary.

Q.2	Long Answers: (Any One)	
	1) Write short note on source of impurities.	10
	2) Write short note on major intra and extra cellular elements.	10
Q.3	Short Answers: (Any Two)	
	1) Write down short note limit test of Iron.	05
	2) Write method of preparation and assay method of sodium fluoride.	05
	3) Derive henderson equation for acidic and basic buffers.	05

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Q.1 Multiple Choice Questions:

- | | | |
|------|---|----|
| (1) | What is buffer capacity equation: - | 01 |
| | (a) $\beta = d[B]/dpH$ (b) $\alpha = d[A]/dpH$ | |
| | (c) $\beta = d[A]/dpH$ (d) $\alpha = d[B]/dpH$ | |
| (2) | Vitamin not require for tooth formation | 01 |
| | (a) B (b) C (c) A (d) D | |
| (3) | Which of the following compound is not use as dentifrices? | 01 |
| | (a) calcium Carbonate (b) calcium carbonate | |
| | (c) calcium gluconate (d) Strontium Chloride | |
| (4) | In limit test of arsenic, arsenic acid converted to | 01 |
| | (a) Arsenious acid (b) Arsine | |
| | (c) stannous acid (d) none | |
| (5) | Lewis acid | 01 |
| | (a) Electron acceptor (b) electron donor | |
| | (c) H ⁺ ion acceptor (d) H ⁺ ion donor | |
| (6) | Acidosis mean | 01 |
| | (a) Decrease level of H ⁺ (b) Increase level of H ⁺ | |
| | (c) Decrease level of co ₂ (d) all of above | |
| (7) | PH range of blood | 01 |
| | (a) 6.45 (b) 7.45 (c) 8.2 (d) 3.2 | |
| (8) | Monograph contain | 01 |
| | (a) Dosage (b) Purity (c) technique (d) storage | |
| (9) | Isotonic solution mean | 01 |
| | (a) same osmotic pressure (b) same concentration | |
| | (c) Both of above (d) none | |
| (10) | What is normal saline solution | 01 |
| | (a) 0.9% KCl (b) 9.0% NaCl (c) 9.0% KCl (d) 0.9% NaCl | |