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## PARUL UNIVERSITY

	PARUL INSTITUTE O	OF PHARMACY	
	B.PHARM FIRST	SEMESTER	
	SECOND INTERNAL EXA	MINATION :2020-21	
Subject Cod Time: 01:15 Instructions 1. Figures to the	pm to 02:30 pm	Date: 13/02/2021 Total Marks: 30	
Q.1 Multi	ple Choice Questions:		
(1)	Which of the following is Aprotic solvent a) Chloroform	t used in Non aqueous Titration ? b) Water	01
	c) Both of above	d) None	
(2)	When AgNO <sub>3</sub> reacts with NaCl, the precip	pitates offorms.	01
(3)	a) AgNCI     c) NaCl  Most widely used titrant in case of Comp	b) AgCl d) NaNO <sub>3</sub>	01
40	<ul><li>a) Ethylene diamine tetrachloride</li><li>c) Ethylene diamine tetraacetic acid</li></ul>	<ul><li>b) Ethylene diamidetereaacetic acid</li><li>d) Ethylene diamine tetra chloride</li></ul>	01
(4)	In case of Assay of calcium gluconate, one must use warm water to dissolve calcium gluconate because		
	a) It is freely soluble in cold water		
	b) It is freely soluble in warm water		
	c) It is freely insoluble in warm water	er	
	d) It is freely insoluble in cold water		
(5) Which of the following can be considered as a key word for co – precipitation in gravimetry?		as a key word for co - precipitation	01
	a) Simultaneous precipitation	b) Precipitation before the actual process	
	c) Precipitation after the actual process	d). None of the above	

(6)	Assay	of CuSO4 is which type of	of titration?		01
	a)	Redox	b)	Precipitation	
	c)	Complexometry	d)	diazotization	
(7)	What i	s limit test?			01
	a)	test for organic impurity	b)	test for inorganic impurity	
	c)	test for purity	d)	none	
(8)	The gla	ass membrane electrode is	responsive t	to	01
	a)	Mg+2	b)	Na+	
	· c)	H+	d)	Al+3	
(9)	How n	nany significant figures are t	here in the nu	imber 0.02040?	01
	(a) 5		(b) 5		
	(c) 7		(d) 1		
(10)	The at	omic weight of silicon is 28	3.0855. Roun	d this to 3 significant figures.	01
	a)	28.09	b)	28.1	
	c)	28.0	d)	28.08	

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# PARUL INSTITUTE OF PHARMACY

### B.PHARM FIRST SEMESTED

	THE SEMESTER		
		SECOND INTERNAL EXAMINATION :2020-21	
	Tin Ins 1. F	Subject Name: Pharmaceutical Analysis Subject Code: BP102T Time: 01:15 pm to 02:30 pm Instructions:  1. Figures to the right indicate full marks. 2. Make suitable assumptions wherever necessary.	
Q.1	Mul	tiple choice questions	10
Q.2	Lon	g Answers: (Any One)	
	(1)	Consider the titration of 1N HCl vs 1N NaOH. The volume of analyte solution (1N HCl) is 100ml. Calculate the pH value of the final solution after addition of 0ml, 7ml, 96ml, 107ml and 112 ml of the Titrant (1N NaOH).	10
	(2)	What different types of electrodes are used in Potentiometry? Explain Any two with construction and working.	10
Q.3	Shor	t answers (Any two)	
	(1)	Write difference between Volhard's method and Modified Volhard's method detail concept.	05
	(2)	Write a note of Diazotization type of titration	05
	(3)	Explain Indicators used in Redox titration.	05

05